Chemistry 115 Name

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Exam 2b October 13, 2010

Multiple Choice (30 points)

Page 5 (24 points)

Page 6 (26 points)

Page 7 (20 points)

Total (100 points)

All work must be shown to receive credit. Give all answers to the correct number of significant figures

Avogadros number = 6.022 x 1023 /mol

Grossmont College

Periodic Table

|  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| IA |  |  |  |  |  |  |  |  |  |  | |  |  |  |  |  | VIIA | NOBLE GASES |
| 1  **H**  1.008 | IIA |  |  |  |  |  |  |  |  |  | |  | IIIA | IVA | VA | VIA | 1  **H**  1.008 | 2  **He**  4.002 |
| 3  **Li**  6.941 | 4  **Be**  9.012 |  |  |  |  |  |  |  |  |  | |  | 5  **B**  10.81 | 6  **C**  12.01 | 7  **N**  14.01 | 8  **O**  16.00 | 9  **F**  19.00 | 10  **Ne**  20.18 |
| 11  **Na**  23.00 | 12  **Mg**  24.30 | IIIB | IVB | VB | VIB | VIIB | VIII VIII VIII | | | | IB | IIB | 13  **Al**  27.00 | 14  **Si**  28.09 | 15  **P**  30.97 | 16  **S**  32.06 | 17  **Cl**  35.45 | 18  **Ar**  39.95 |
| 19  **K**  39.10 | 20  **Ca**  40.08 | 21  **Sc**  44.96 | 22  **Ti**  47.90 | 23  **V**  50.94 | 24  **Cr**  52.00 | 25  **Mn**  54.94 | 26  **Fe**  55.85 | 27  **Co**  58.93 | 28  **Ni**  58.70 | | 29  **Cu**  63.55 | 30  **Zn**  65.38 | 31  **Ga**  69.72 | 32  **Ge**  72.59 | 33  **As**  74.92 | 34  **Se**  78.96 | 35  **Br**  79.90 | 36  **Kr**  83.80 |
| 37  **Rb**  85.47 | 38  **Sr**  87.62 | 39  **Y**  88.91 | 40  **Zr**  91.22 | 41  **Nb**  92.91 | 42  **Mo**  95.94 | 43  **Tc**  (99) | 44  **Ru**  101.1 | 45  **Rh**  102.9 | 46  **Pd**  106.4 | 47  **Ag**  107.9 | | 48  **Cd**  112.4 | 49  **In**  114.8 | 50  **Sn**  118.7 | 51  **Sb**  121.8 | 52  **Te**  127.6 | 53  **I**  126.9 | 54  **Xe**  131.3 |
| 55  **Cs**  132.9 | 56  **Ba**  137.3 | 57  **La**  138.9 | 72  **Hf**  178.5 | 73  **Ta**  180.9 | 74  **W**  183.9 | 75  **Re**  186.2 | 76  **Os**  190.2 | 77  **Ir**  192.2 | 78  **Pt**  195.1 | 79  **Au**  197.0 | | 80  **Hg**  200.6 | 81  **Tl**  204.4 | 82  **Pb**  207.2 | 83  **Bi**  209.0 | 84  **Po**  (209) | 85  **At**  (210) | 86  **Rn**  (222) |
| 87  **Fr**  (223) | 88  **Ra**  226.0 | 89  **Ac**  227.0 | 104  **Rf**  (261) | 105  **Db**  (262) | 106  **Sg**  (263) | 107  **Bh**  (262) | 108  **Hs**  (265) | 109  **Mt**  (266) | 110  **??**  (269) |  | |  |  |  |  |  |  |  |

|  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| 58  **Ce**  140.1 | 59  **Pr**  140.9 | 60  **Nd**  144.2 | 61  **Pm**  (147) | 62  **Sm**  150.4 | 63  **Eu**  152.0 | 64  **Gd**  157.3 | 65  **Tb**  158.9 | 66  **Dy**  162.5 | 67  **Ho**  164.9 | 68  **Er**  167.3 | 69  **Tm**  168.9 | 70  **Yb**  173.0 | 71  **Lu**  175.0 |
| 90  **Th**  232.0 | 91  **Pa**  231.0 | 92  **U**  238.0 | 93  **Np**  (237) | 94  **Pu**  (244) | 95  **Am**  (243) | 96  **Cm**  (247) | 97  **Bk**  (247) | 98  **Cf**  (251) | 99  **Es**  (252) | 100  **Fm**  (257) | 101  **Md**  (258) | 102  **No**  (259) | 103  **Lr**  (260) |

Lanthanide series

Actinide series

Part I – Multiple Choice (30 points)

Exam 1 multiple choice questions

1. Atomic emission spectra are due to electrons
   1. being removed from an atom.
   2. in an atom dropping from one energy level to a lower one.
   3. in an atom rising from one energy level to a higher one.
   4. being added to an atom.
   5. changing state from solid to liquid.
2. According to the Pauli exclusion principle, any orbital can hold at most *\_\_\_\_\_\_\_\_* electrons.
   1. 6
   2. 8
   3. 2
   4. 18
   5. 10
3. The electron configuration of an atom shows
   1. the number of electrons in each electron energy level.
   2. the number of isotopes possible.
   3. a description of the shape of each electron energy level.
   4. a diagram of an atomic nucleus.
   5. the maximum number of electrons each electron energy level can hold.
4. The atomic radius of potassium is smaller than the atomic radius of *\_\_\_\_\_\_\_\_.*
   1. lithium
   2. fluorine
   3. hydrogen
   4. cesium
   5. sodium
5. Valence electrons are electrons located
   1. in the nucleus of an atom.
   2. in the first energy level of an atom.
   3. in the outermost energy level of an atom.
   4. throughout the atom.
   5. in the first three energy levels of an atom.
6. The ionization energy of chlorine is lower than the ionization energy of *\_\_\_\_\_\_\_\_.*
   1. sodium
   2. hydrogen
   3. lithium
   4. calcium
   5. fluorine
7. To form an ion, a sodium atom
   1. gains one electron.
   2. gains two electrons.
   3. loses one electron.
   4. loses seven electrons.
   5. loses two electrons.
8. An anion always
   1. has a negative charge.
   2. has a positive charge.
   3. contains a group of two or more atoms with a positive charge.
   4. contains a metal and a nonmetal.
   5. forms covalent bonds.
9. Which one of the following elements forms two or more ions with different ionic charges?
   1. K
   2. F
   3. Ca
   4. O
   5. Fe
10. A group of covalently bonded atoms that has an overall electrical charge is called a(n) \_\_\_\_\_\_\_\_.
    1. ionic compound
    2. polyatomic ion
    3. anion
    4. cation
    5. molecule
11. According to the IUPAC nomenclature system, the types of compound that use prefixes in their names are \_\_\_\_\_\_\_\_.
    1. ionic compounds
    2. ionic compounds involving transition metals
    3. polyatomic ions
    4. covalent compounds
    5. compounds that contain polyatomic ions
12. What is the name of this compound?

CH3- CH2- CH2- CH2- CH2- CH2- CH3

* 1. hexane
  2. octane
  3. butane
  4. heptane
  5. pentane

1. Avogadro's number is the number of
   1. particles in 1 mol of a substance.
   2. amu in 1 mol of a substance.
   3. grams in 1 mol of a substance.
   4. moles in 6.022 × 1023 grams of an element.
   5. moles in 6.022 × 1023 amu of an element.
2. A chemical equation is balanced when
   1. the total number of molecules is the same in reactants and products.
   2. the total number of ions is the same in reactants and products.
   3. the sum of the coefficients of the reactants is equal to the sum of the coefficients of the products.
   4. the charge on each atom is the same in reactants and products.
   5. the number of atoms of each element is the same in reactants and products.
3. The reaction of carbon dioxide to form carbon monoxide and oxygen is what type of reaction?

2CO2 (*g*) → 2CO (*g*) + O2 (*g*)

* 1. single replacement
  2. combination
  3. oxidation
  4. decomposition
  5. double replacement

Part 2 – Problems and Short Answer (70 points)

1. (3 points) What is an orbital?
2. (3 points) Write the complete electron configuration of sulfur.
3. (3 points) Write the shorthand electron configuration of Technetium (Tc)
4. (5 points) Describe how atomic size changes as you move across the periodic table to the right and explain the reason for this trend.
5. (4 points) Circle the element with the higher ionization energy

Calcium or Barium?

Potassium or Arsenic?

1. (6 points) Name the following compounds
   1. BaSO4
   2. NH4Br
   3. K3P
   4. CCl4
2. (6 points) Give the correct formula for the following compounds
   1. Titanium(II) nitride
   2. Aluminum hydroxide
   3. Tribromine octafluoride
   4. Cadmium oxide
3. (3 points) Calculate the mass of 2.99 moles of gold (Au)
4. (4 points) Calculate the number of atoms of gold in 5.31 moles of gold.
5. (4 points) Calculate the molar mass of benzaldehyde (C7H6O)
6. (4 points) Calculate the number of moles of benzaldehyde in 76.1 grams of benzaldehyde.
7. (5 points) Calculate the number of atoms of carbon in 6.00 g of benzaldehyde.
8. (5 points) Determine the empirical formula of methyl butyrate, the flavor of apples. It is composed of 58.80% C, 9.87% H, and 31.33% O.
9. (3 points) A compound has an empirical formula of C6H7N and a molar mass of 570.8 g/mol. What is the molecular formula of the compound?
10. (6 points) Balance the following equations
    1. FeS + HCl 🡪 FeCl2 + H2S
    2. C6H14 + O2 🡪 CO2 + H2O
11. (6 points) Match each of the following molecules to the correct functional group.
    1. 
    2. 
    3. 
    4. 
    5. 
    6. 

Alcohol

Aldehyde

Alkene

Alkyne

Amide

Amine

Aromatic

Carboxylic acid

Ester

Ether

Ketone